



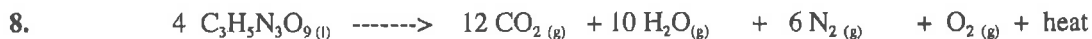
a. If 56.3 grams of barium chloride react with excess sodium phosphate, how many grams of barium phosphate will form?

b. Learn the formula for % yield: $\% \text{ yield} = \frac{\text{actual (lab) value}}{\text{expected (stoichiometry) value}} \times 100\%$

c. Suppose that when the reaction in (a) is done in lab, only 53.3 grams of barium phosphate precipitate are actually collected. What was the percent yield for the reaction?

d. What mass of sodium phosphate is needed to produce 100.0 grams of sodium chloride in the reaction?

e. If 1.0 moles of sodium phosphate react, how many moles of barium phosphate will form?



a. What mass of nitroglycerine are needed to produce 50.0 grams of oxygen gas?

b. If 450 grams of nitroglycerine decompose, how many moles of carbon dioxide gas will be produced?



a. If 1.0×10^{22} copper atoms react, what mass of NO gas can form?

b. Suppose that the actual mass of NO collected in (a) is only 0.31 grams. Calculate the % yield.

c. If 10.0 grams of copper react, what mass of water can form in the reaction?