

Aluminum and copper (II) chloride reaction.

Purpose: To observe a chemical reaction

Equation for the reaction:



Procedure. (In addition to the equipment mentioned below, you may use any of the lab instruments in your drawer or on the counter to help with observations/measurements. Think about which quantitative data you might want to record before you begin.)

1. Fill a 100 mL beaker about ¼ full of tap water. Add about 1 level spoonful of copper (II) chloride to the water. Allow the system of water and crystals to stand for a few minutes and record any qualitative and quantitative observations.
2. Stir the system until the crystals dissolve. Record observations.
3. Loosely roll up a square of aluminum foil, and add the foil to the beaker/contents. Record observations.

Qualitative Data: (no numbers used)

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Quantitative Data: (measurements including numbers and units)

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Data Analysis:

1. Review the following terms (see pages 15-18 as needed): Element, compound, homogeneous mixture, heterogeneous mixture. Classify each of the following substances as an element, compound, homogeneous mixture, or heterogeneous mixture.

Water: \_\_\_\_\_

Copper (II) chloride: \_\_\_\_\_

Copper (II) chloride and water, after dissolving together: \_\_\_\_\_

Copper (II) chloride and water, when in layers (before dissolving): \_\_\_\_\_

Dark green liquid with brown chunks of solid and fizz: \_\_\_\_\_

Aluminum foil (before the reaction.) \_\_\_\_\_

