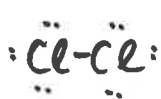


Write the total number of valence electrons next to each molecule or ion below. Then draw the Lewis Dot Structure! Assume that all of the elements will follow the octet rule, except H, which will bond to get 2 valence electrons.

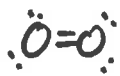
Cl₂ 14



F₂ 14



O₂ 12



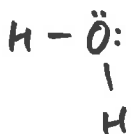
N₂ 10



H₂ 2



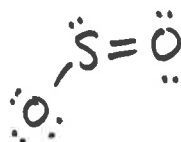
H₂O 8



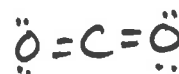
SCl₂ 20



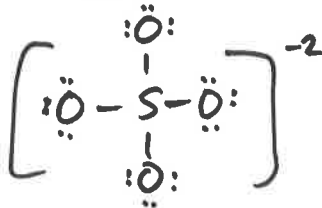
SO₂ 18



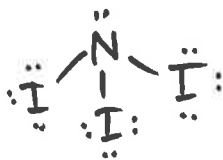
CO₂ - 16



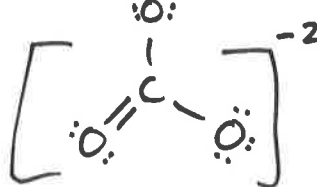
SO₄⁻² 32



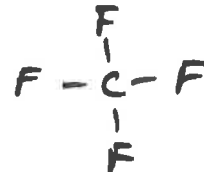
NI₃ 26



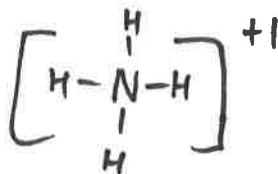
CO₃⁻² 24



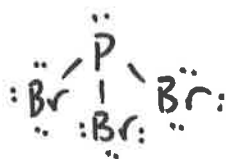
CF₄ 32



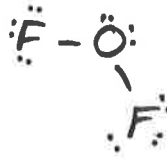
NH₄⁺¹ 8



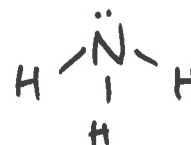
PBr₃ 26



OF₂ 20



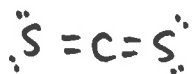
NH₃ - 8



HCN 10



CS₂ 16



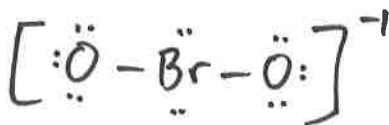
CO 10



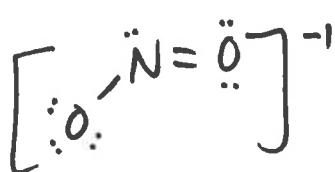
HF - 8



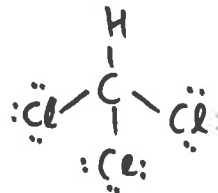
BrO₂⁻¹ 20



NO₂⁻¹ 18



CHCl₃ 26



CF₂Cl₂ 32

