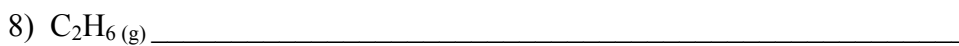


Write equations showing the **heat of formation** of the following compounds. Be sure to include the states of matter and include the energy term (kJ) in the equation, on the correct side. Some equations may need fractional coefficients to end up with one mole of product. Use your thermochemical reference handout.

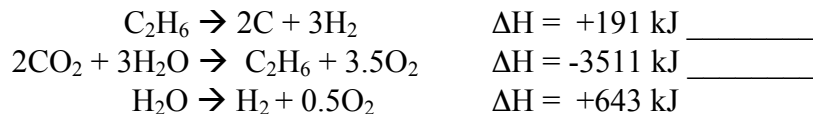


Write equations showing the heat of combustion of the following substances. Be sure to include states of matter and include the energy term (kJ) in the equation, on the correct side. Some equations may need fractional coefficients to end up with one mole of fuel. Use your thermochemical reference HO.



HESS'S LAW REVIEW MASTER EQ: CO₂ \rightarrow C + O₂ $\Delta H =$ _____ ????

Given:



Rewrite:

\rightarrow
 \rightarrow
 \rightarrow

Flip over for bookwork:

Bookwork: page 522 (#1 and #3)

page 531 (#4 and #6)
