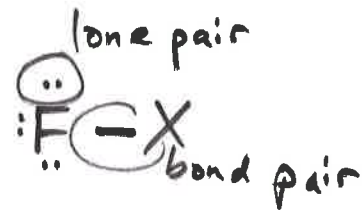


3/17 Notes : Covalent Bonding

Lewis Structures

- ① count up all ve^-
- ② Find central atom (least EN)
- ③ Try all single bonds, lone pairs
- ④ See if $ve^- \neq$ matches
- ⑤ If no, try double or triple bond.

$$\frac{H_2 O}{2 + 6} = 8$$



H-

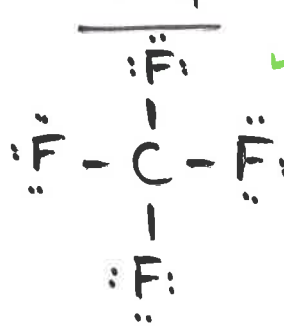
ex) $I_2 - 14$



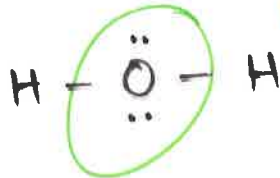
$HF = 8$



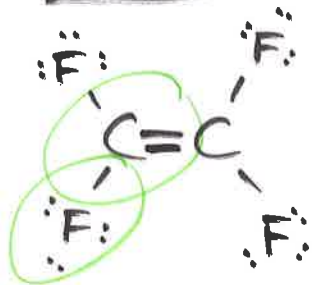
$CF_4 = 32$



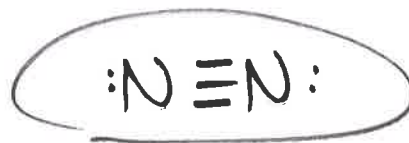
$H_2O = 8$



$C_2F_4 = 36$



$N_2 = 10$



10 ✓



Generalities

C	group like	4 bonds,	0	Lone pairs
N		3 bonds,	1	
O		2 bonds,	2	
F		1 bond,	3	