

Notes 9/26 % Comp continued

1) You extract Zn from $ZnCl_2$ in lab.

DATA:

5.00g $ZnCl_2$ total

2.20g Zn extracted

LAB % Zn in $ZnCl_2$

$$\% Zn = \frac{2.20g Zn}{5.00g ZnCl_2} \times 100 = \frac{44.0\% Zn}{LAB}$$

THEO % Zn in $ZnCl_2$

$$\% Zn = \frac{Zn}{ZnCl_2} = \frac{65.41}{65.41 + 2(35.45)} \times 100 = 47.99\%$$

$$\% \text{ error} = \frac{|T-L|}{T} \times 100 = \frac{Zn}{THEO}$$

$$\frac{|47.99 - 44.0|}{47.99} \times 100 = \boxed{8.31\% \text{ error}}$$